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three types of hydrogen bonds form and exist in the water molecule. one or more hydrogen atoms are directly attached to an oxygen atom. when oxygen is involved, we refer to this as an oxygen-hydrogen bond. each oxygen-hydrogen bond is actually between a water molecule, which we refer to as an oxygen atom, and a hydrogen atom, which we refer to as a hydrogen atom. the oxygen is called the negative pole and the hydrogen is called the positive pole. one oxygen-hydrogen bond forms when the two atoms are close together. another type of bond is formed when the hydrogen atoms are close together but the oxygen atom is not close enough for there to be an oxygen-hydrogen bond. in the water molecule, three oxygen atoms and three hydrogen atoms are involved in forming a bond between the two atoms.

each oxygen atom forms one bond with a hydrogen atom and each hydrogen atom forms one bond with an oxygen atom. each oxygen-hydrogen bond is said to form between the oxygen and hydrogen atoms. the third type of bond is said to form between the water molecules. the hydrogen and oxygen atoms are similar to dots on a piece of paper so that a 3-dimensional view of the water molecule is shown. this is our image of the water molecule. three organic compounds named alkanes, alkenes, and alkynes are found in nature. they are hydrocarbon molecules that consist of carbon-hydrogen bonds. the alkanes are also called methanes, the alkenes are also called ethene, and the alkynes are also called ethylenes. methane is called CH_4 and ethylene is called C_2H_4 .

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